**Notes: Mixtures and Solutions**

**Define Mixture**

**Define Heterogeneous**

**Example: \_\_\_cereal, salad\_\_\_\_\_\_\_\_\_\_\_\_**

**Define Homogeneous**

**Example: \_\_milk, honey, metal\_\_\_\_**

**3 types of Mixtures**

**True solution ex. Ex. Lemonade, sugar water**

**Colloid ex. . Ex. Smoke, jelly, milk, mayo**

**Suspension ex. . Ex. Muddy water, Jello.**

**Define Solution**

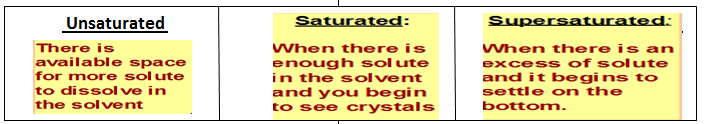
**Determined by: particle size of dissolved substance**

**Define Solute:**

**Example: \_\_\_\_sugar\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Define Solvent**

**Example: \_\_\_\_\_water\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**



**Define Solubility**

**What does a Solubility Graph tell us?**

**Define Compound**

**Define Element**

**2 or more substance put together but still remains separate**

**Not mixed well. Can see separate particles**

**Mixed well. Cannot see individual particles**

**Description: Mixture of 2 or more substances w/dissolving.**

**Light passes through**

**Description: A solution that contains particles that do not settle ever. Light does not pass through**

**Description: heterogeneous mixture with a fluid and particles. The particles will settle out.**

**Homogeneous mixture. One substance dissolves in another**

**In a solution, the substance that is dissolved**

**In a solution, the substance that does the dissolving**

**The amount of solute a solvent can dissolve at a particular temperature**

**It explains & compares the amount of different solutes dissolved at a particular temperature.**

**Not a Mixture. It’s a combination of 2 or more elements chemically combined making a new substance entirely.**

**A substance that cannot be broken down. Defined by the number of protons it has.**